EFFECTS OF SALT CONCENTRATION ON VAPOR - LIQUID EQUILIBRIUM (VLE) OF AZEOTROPIC MIXTURE IN ULTRASONIC DISTILLATION SYSTEM

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UNIVERSITI TEKNOLOGI MALAYSIA

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BORANG PENGESAHAN STATUS TESIS ♦

JUDUL : **EFFECTS OF SALT CONCENTRATION ON VAPOR-LIQUID EQUILIBRIUM (VLE) OF AZEOTROPIC MIXTURE IN ULTRASONIC DISTILLATION SYSTEM**

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NOR FAIZAH BINTI RAZALI

A thesis submitted as fulfillment of the award of the degree of Master of Engineering (Chemical)

Faculty of Chemical and Natural Resources Engineering
Universiti Teknologi Malaysia

NOVEMBER 2008

"I declare that this thesis entitled "**EFFECTS OF SALT CONCENTRATION ON VAPOR-LIQUID EQUILIBRIUM (VLE) OF AZEOTROPIC MIXTURE IN ULTRASONIC DISTILLATION SYSTEM**" is the result of my own work except cited in references. The thesis has not been accepted for any degree and is not concurrently submitted in candidature of any degree."

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Specially dedicated to

my beloved father, Razali bin Ismail, my beloved mother, Rujomah bte Jamil

and

those people who have guided and inspired me throughout my journey of education

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ABSTRACT

Methanol and methyl acetate form an azeotrope in their mixtures at 34.78 mole % of methanol. It is difficult and may be impossible to separate azeotropic mixture using conventional distillation column. This phenomenon gives big challenges in the chemical industry in order to solve it. In this study, potassium chloride was added in ultrasonic distillation system to study the combination effect of salt and ultrasonic wave on methanol-methyl acetate mixture. The salt was added with different range of concentration to select the optimum concentration that can eliminate the azeotropic point. Ultrasonic wave with frequency of 25 kHz and intensity of 200 W/A.cm^2 were used. The studies on the effect of different salt concentration at 0 wt%, 5 wt%, 10 wt% and 15 wt% to VLE of binary mixtures were done at that frequency and intensity to obtain the best salt concentration. The results obtained show that, as the salt concentration increased in the liquid phase, the equilibrium line shifts upwards and in the same time, the azeotropic point also move upward. The salt concentration used in this work give the results in the following order 15 wt% >10 wt% > 5 wt% > 0 wt% where the azeotropes point form at 70 mole %, 54 mole %, 48 mole % and 38 mole % of methanol accordingly. As the result, the best concentration of the potassium chloride for the methanol-methyl acetate separation in this project was at 15 wt% of concentration. These results show that the combination of ultrasonic and salt as a separating agent gave positive results and have a potential to be apply for industry in the future.

ABSTRAK

Campuran metanol dan metil acetat membentuk azeotrop pada titik 34.78 mol % dari metanol. Sebatian azeotrop ini merupakan campuran yang agak sukar dipisahkan dan mungkin tidak boleh dipisahkan oleh penyulingan biasa. Hal ini membuatkan industri kimia pada masa kini menghadapi cabaran yang agak besar untuk mengatasi masalah ini. Kajian ini memperkenalkan kaedah baru dalam percubaan untuk memisahkan campuran azeotrop ini, dimana garam digunakan sebagai agen pemisahan dalam sistem penyulingan ultrabunyi. Garam yang digunakan adalah kalium klorida, dimana garam ini di campurkan ke dalam sebatian metanol-metil acetat, dalam julat kepekatan yang berbeza agar nilai optimum kepekatan garam dapat diperoleh. Frekuensi gelombang ultrabunyi yang digunakan adalah 25 kHz, manakala keamatan ultrabunyi yang dibekalkan adalah pada 200 $W/A.cm^2$. Julat kepekatan garam yang digunakan adalah sebanyak 0 wt%, 5 wt%, 10 wt% dan 15 wt%. Keputusan yang diperoleh menunjukkan kesan yang agak baik apabila lengkungan keseimbangan metanol-metil acetat berganjak menjauhi garisan $45^{\circ}C$ dan seterusnya melonjakkan nilai titik azeotrop. Hasil yang diperoleh adalah mengikut urutan 15 wt% >10 wt% > 5 wt% > 0 wt% dimana titik azeotrop yang diperoleh adalah pada 70 mol %, 54 mol %, 48 mol % dan 38 mol % dari metanol mengikut urutan. Kesimpulannya, kepekatan garam kalium klorida yang terbaik dalam projek ini adalah pada kepekatan 15 wt %. Keputusan ini menunjukkan kaedah yang diperkenalkan ini memberi hasil yang positif. Ini memungkinkan kaedah ini untuk diteruskan kajiannya, seterusnya diaplikasikan dalam industri pada masa hadapan.

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LIST OF SYMBOLS AND ABBREVIATION

VLE	-	Vapor-Liquid Equilibrium
m	-	Meter
°C	-	Degree Celsius
ml	-	Mill liter
g	-	Gram
KCl	-	Potassium Chloride
%	-	Percentage

CHAPTER I

INRODUCTION

1.1 Introduction

Distillation is the most widely applied separation technology and will continue as important process in the future because there is simply no industrially viable alternative around. Eventhough this technique confronted challenges from other technologies, it still improves from time to time and moves to a higher level of sophistication. Nevertheless, there are still many technical barriers faced by distillation. Not all liquid mixture can be separated by ordinary fractional distillation. The separation becomes difficult and expensive when the components of the system have low relative volatilities ($1.00 < \alpha < 1.05$). This is because a large numbers of trays are required and, usually, a high reflux ratio as well. A different problem occurs if the system forms azeotropes, where the azeotropic composition limits the separation (Cheresources.com website, 2008).

Azeotrope means literally that the vapor boiling from a liquid has the same composition as the liquid. The azeotropic mixture depends upon the degree of non-ideality of a mixture and the difference in boiling points between the two pure components (Kim et al., 1997).

It has been known that the separation of components by simple distillation is impossible if the mixture exhibits an azeotropic at a specified temperature and pressure. For a mixture that has an azeotrope, the separation process relies on the addition of specially chosen chemicals to eliminate the azeotropes entirely. Salts are one of the mass-separating agents that can be introduced in distillation system in order to solve this problem (Banat et al., 1997).

A salt dissolved in a mixture of volatile components may affect the activities of the components through the formation of complexes (Yao et al., 1999). The salt will dissociates into ions in the liquid mixture and alters the relative volatilities to make the separation becomes possible (Banat et al., 1997).

Normally, even little salt may bring an appreciable effect on the relative volatility where this is called the effect of preferential salvation. The use of a salt instead of another separating agent in distillation gives several advantages including lower energy consumption. It also gives high purity of the overhead products because salts are non-volatile and hence do not evaporate or condense during distillation process (Banat et al., 1997). Salts also has lower toxicity lever comparing with other liquid separation agents such as benzene.

The previous research has verified the feasibility of using ultrasonic wave to enhance the separation of binary mixtures in distillation column and overcoming the challenges of azeotropic separation. In this project, salt was used to be tested in ultrasonic distillation system. The salt was introduced in the flask with different concentration. Once the optimum salt concentration was achieved, further analysis can be done by calculate its relative volatility.

This research was about to see the effect of salts in ultrasonic distillation system in order to enhance the separation process. The selected binary mixture was methanol and methyl acetate, while the salt used was potassium chloride (KCl).

1.2 Problem statement

Distillation is the common separating method to separate components in liquid mixtures. However, this process may be complicated by the formation of azeotropes due to non idealities in the mixture. These azeotropes can make a given separation impossible by conventional distillation processes. Industries always looking forward to obtain the solution for this problem.

Salt is one of the separating agents in distillation process for separating close-boiling or azeotropes systems that cannot easily be purified using ordinary distillation. Salt has been proved for eliminating the azeotropes entirely. In this work, vapor-liquid equilibrium (VLE) studies were conducted to determine the optimum concentration of selected salt in ultrasonic distillation system.

1.3 Objective

The main objective of this research is to study the separation of azeotropic mixture by using salt on VLE in ultrasonic distillation system.

1.4 Scope of Research

The scope of this research was to identify the appropriate salt concentration to be applied in ultrasonic distillation system.

CHAPTER II

LITERATURE REVIEW

2.1 Distillation

Distillation is a method of separating mixtures where a liquid or vapor mixture of two or more substances is separated into its component fractions of desired purity, by the application and removal of heat. The separation of mixture is based on differences in their volatilities in a boiling liquid mixture (Fair, 2000). Distillation is the most common separation technique and it consumes enormous amounts of energy, both in terms of cooling and heating requirements. It can contribute to more than 50% of plant operating costs. The best way to reduce operating costs of existing unit is to improve their efficiency and operation via process optimization and control (Distillation website, 2008).

Distillation exists either in batch or continuous mode. In batch distillation, the composition of the source material, the vapors of the distilling compounds and the distillate change during the distillation. In batch distillation, a still is charged (supplied) with a batch of feed mixture, which is then separated into its component fractions which are collected sequentially from most volatile to less volatile, with the bottoms (remaining least or non-volatile fraction) removed at the end. The still can then be recharged and the process repeated. In continuous distillation, the source materials, vapors, and distillate are kept at a constant composition by carefully adding the source material and removing fractions from both vapor and liquid in the system. This results in a better control of the separation process (Wikipedia website, 2008).

2.2 Vapour-Liquid-Equilibrium (VLE)

2.2.1 Introduction

VLE measurements are tedious and time-consuming because measurement conditions are often controlled and recorded manually. Cost reduction can be achieved by affordable automation, which permits a more efficient operation of the apparatus and, in some cases, an increase in accuracy. One problem associated with automation is that researchers working with experimental thermodynamics seldom seem to have the expertise needed in laboratory automation (Ussi-Kyyny, 2004).

However, when automation expertise has been successfully created in the laboratory, the goal should be to implement data acquisition programs and automation software to increase the measurement output of the experimental devices. It is thereby possible to decrease the cost of one individual measurement point substantially. Suitable methods for determination of VLE vary. In some cases several methods can be applied, but in the most difficult cases measurements are almost

impossible. The selection of methods and apparatuses depend on the physical properties of the system studied such as vapor pressure, component stability, material compatibility, measurement accuracy and safety. The properties determined specifically for binary vapor liquid equilibrium systems are temperature, pressure and the compositions of the constituent phases (Ussi-Kyyny, 2004).

The determination of composition is the most complex task.The devices needed are often expensive and there is no universal analytical device that is suitable for all components. Gas chromatography is used most often for the determination of the composition of phases. Other methods for composition determination, although seldom applied in VLE measurements, include mass spectrometry, various spectroscopic methods, and density and refractive measurement (Ussi-Kyyny, 2004).

2.2.2 Vapour-Liquid-Equilibrium (VLE) Curves

Constant pressure VLE data is obtained from boiling point diagrams. Figure 2.1 shows the plot that often presented for VLE data of binary mixtures. The VLE plot expresses the bubble-point and the dew-point of a binary mixture at constant pressure. The curved line is called the equilibrium line and describes the compositions of the liquid and vapor in equilibrium at some fixed pressure. This particular VLE plot shows a binary mixture that has a uniform vapor-liquid equilibrium that is relatively easy to separate (Distillation website, 2008).

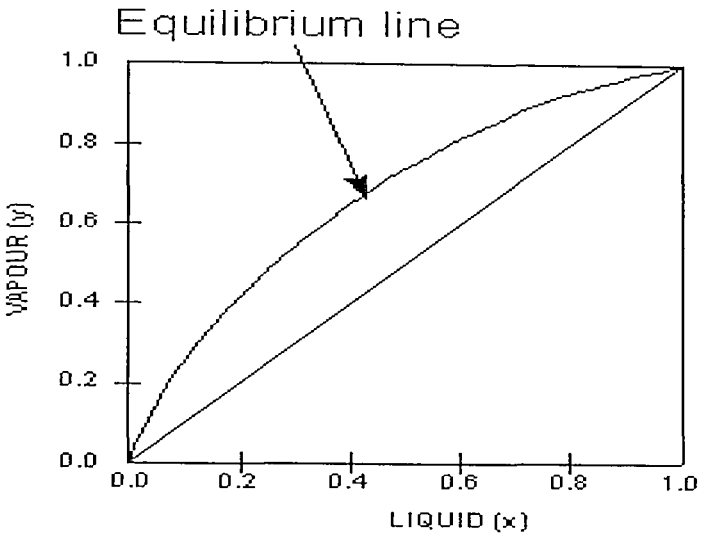

Figure 2.1: VLE graph of binary mixtures

2.3 Azeotropic Formation in Binary Mixtures

The most intriguing VLE curves are generated by azeotropic systems. An azeotrope is a liquid mixture which when vaporised, produces the same composition as the liquid. Figure 2.2 shows two different azeotropic systems, one with a minimum boiling point and one with a maximum boiling point. In both plots, the equilibrium curves cross the diagonal lines, and this are azeotropic points where the azeotropes occur. In other words azeotropic systems give rise to VLE plots where the equilibrium curves crosses the diagonals (Distillation website, 2008).

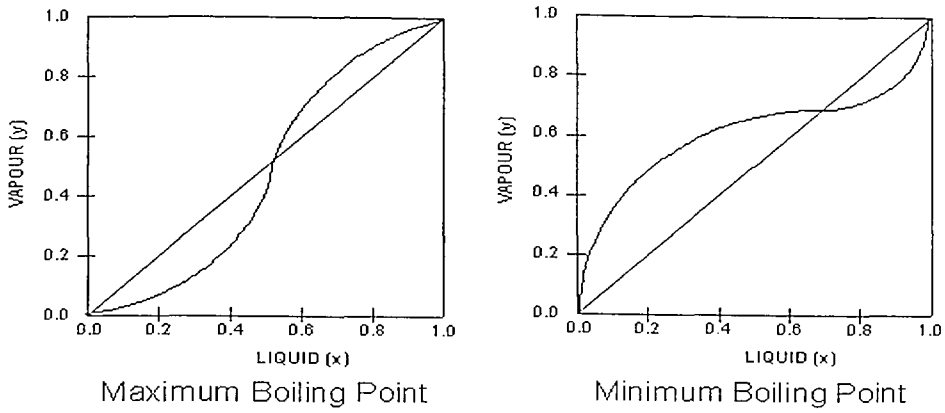

Figure 2.2: Two different azeotropic systems (homogenous azeotrope)

Note the shapes of the respective equilibrium lines in relation to the diagonal lines that bisect the VLE plots. Both plots are however, obtained from homogenous azeotropic systems. Homogenous azeotrope is an azeotrope that contains one liquid phase in contact with vapour. A homogenous azeotrope cannot be separated by conventional distillation. However, vacumn distillation may be used as the lower pressures can shift the azeotropic point. Figure 2.3 shows the VLE curve that also generated by an azeotropic system, in this case a heterogenous azeotrope. Heterogenous azeotropes can be identified by the flat portion on the equilibrium diagram. They may be separated in 2 distillation columns since these substances usually form two liquid phases with widely differing compositions. The phases may be separated using settling tanks under appropriate conditions (Distillation website, 2008).

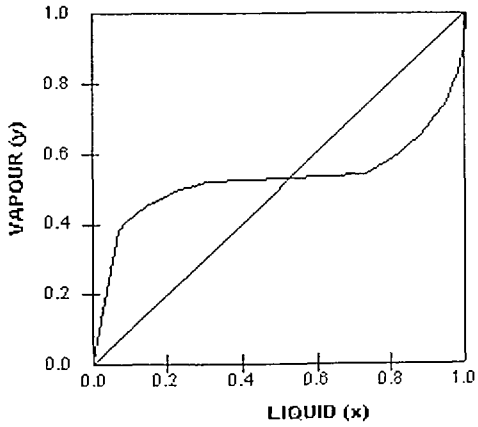

Figure 2.3: Heterogenous azeotrope

2.3.1 Physical Phenomenon Leading to Azeotropy

The tendency of mixture to form azeotrope depends on two factors, which are: (i) the difference in the pure component boiling points, and (ii) the degree of non-ideality (Kim and Simmrock, 1997). The closer the boiling points of the pure components and the less-ideal mixture, the greater the likelihood of an azeotrope The mixtures that have wide differences in their components boiling point may not exhibit an azeotrope even though they form a nonideal mixture.

Most mixtures of organic compounds form nonideal systems. The presence of some specific groups, particularly polar groups such as oxygen, nitrogen, chlorine, and fluorine often results in the formation of azeotropes. The nonideality of mixture depends largely on intermolecular forces of attraction among the mixture components such as dispersion forces, dipole-dipole interactions, dipole-induced dipole interactions, and hydrogen bonding (Swietoslawski, 1963).

2.4 Ultrasound

2.4.1 About Ultrasound

Ultrasonic waves are mechanical pressure waves formed by actuating the ultrasonic transducers with high frequency, high voltage current generated by electronic oscillators. A typical industrial high power generator produces ultrasonic frequencies ranging from 20 - 120 kHz. Typical transducers are normally immersed in the liquid. The generated ultrasonic waves propagate perpendicularly to the resonating surface (Awad, 1996).

The waves interact with liquid media to generate cavitation implosions. High intensity ultrasonic waves create micro vapor bubbles in the liquid medium, which grow to maximum sizes proportional to the applied ultrasonic frequency and then implode, releasing their energies. The higher the frequency, the smaller the cavitation size. The high intensity ultrasonics can also grow cavities to a maximum in the course of a single cycle (Awad, 1996).

2.4.2 Cavitation

Cavitation is the formation and collapse of either gas or vapor bubbles in a liquid subjected to pressure changes. The formation of cavities in liquids is analogous to tensile failure in solids. When the "tensile strength" of a liquid is exceeded, cavities form. Actual values of these "strengths" are much lower than theoretical values, as a result of imperfections (gas pockets) in the liquid which serve as nuclei for cavitation. These nuclei grow, through net diffusion of dissolved gas from the liquid to the nuclei, to form cavitational bubbles. When a high enough pressure amplitude (cavitation threshold) is reached, the nucleus becomes unstable and rapidly grows into a mostly vapor-filled bubble or transient cavity (Busnaina et al., 1994).

2.5 Salts

An additional substance may add to shift the azeotropic point to a more favourable position (Distillation website, 2008). Salt is one of the additional substances that can be added to enhance the separation process. The salt acts as a separating agent by raising the relative volatility of the mixture and by breaking any azeotropes that may otherwise form.

The addition of a salt, instead of a liquid, as a separating agent in distillation provides several advantages including lower energy consumption and high purity of the overhead products because salts are nonvolatile and hence do not evaporate or condense during distillation. The salt effect is believed to be a complex function of salt and solvent interaction and self-interaction among all other components. In most cases, but not necessarily always, the molecules of the more polar solvent, in which the salt is more soluble, are preferentially attracted by the electrostatic field of the salt ions and thereby the molecules of the less polar solvent are liberated to the vapor phase (salted-out) (Banat et al., 1997). The list below show the general advantages of using salt in distillation column:

- Allows continuous operation because of the high efficiency and the low waste of solvent.
- A high purity product can be obtained.
- The relative volatility of mixture is increased, makes the separation processes easier.
- Improves the solvent performance. Compared with normal extractive distillation, the quantity of the solvent to recycle is reduced to its fourth or fifth, the number of theoretical stages required can be reduced to its third, as well as energy consumption.

CHAPTER III

METHODOLOGY

3.1 Introduction

This chapter discussed about the research methodology that applied in this study. A set of distillation apparatus equipped with ultrasonic wave generating equipment was set up as the experimental rig to obtain the vapor-liquid equilibrium data. Only one type of binary mixture was studied in this research. The ultrasonic wave generating equipments, which were ultrasonic transducers and ultrasonic generators, were supplied by Crest Ultrasonic (M) Sdn. Bhd. based in Penang.

3.2 Equipment Set-Up

The VLE data in this study were collected by using a laboratory scale distillation column called Ultrasonic-Distillation System. Figure 3.1 shows the schematic diagram of Ultrasonic-Distillation System. The designed equipment comprises of distillation flask, condenser, thermocouples, heating bath, chiller and ultrasonic generating equipment. The liquid and vapor temperature in the distillation column were measured using ordinary termometer. The prepared binary mixture was placed inside the distillation flask. A heating bath which equipped with temperature controller was used to provide heat for boiling process is shown in figure 3.2 (a), while figure 3.2 (b) shows the chiller that used to provide cooling water for condenser. When the liquid mixture was boiled, the vapor produced will ascend the distillation column and condensed to form liquid droplets that were collected through a valve.

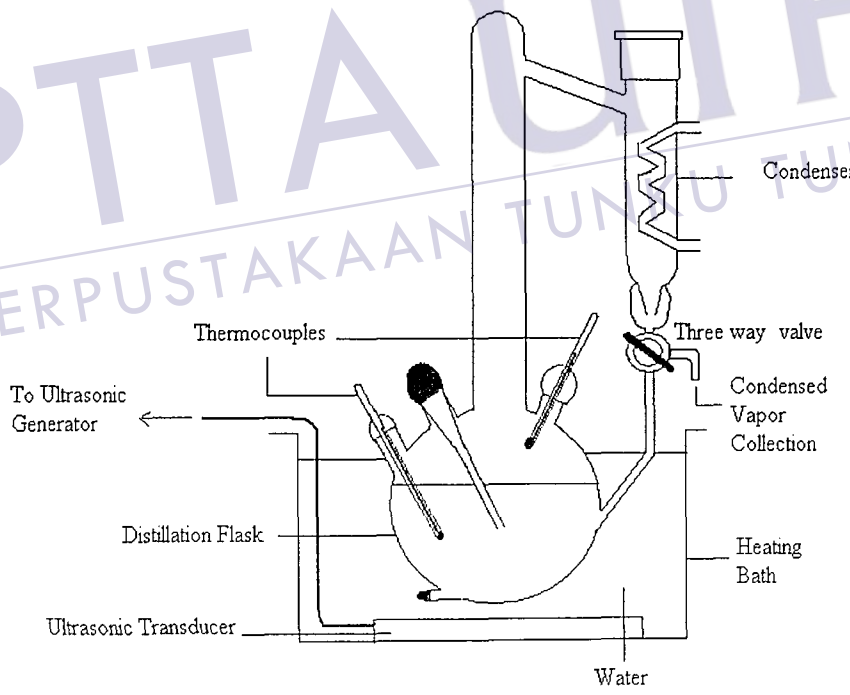

Figure 3.1: Schematic diagram of Ultrasonic-Distillation System

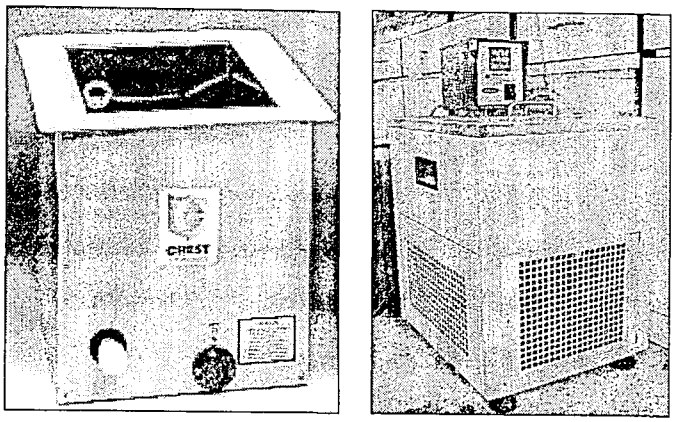

Figure 3.2: (a) Heating bath (b) Chiller

3.2.1 Ultrasonic Wave Generating Equipments

3.2.1.1 Ultrasonic Generator

Figure 3.3 shows the ultrasonic generator that used in this research. It was constructed and supplied by Crest Ultrasonic (M) Sdn. Bhd., Penang. On the front panel of the generator is the "On/Off" switch. The output power intensity is adjustable via a control knob located on the front panel.

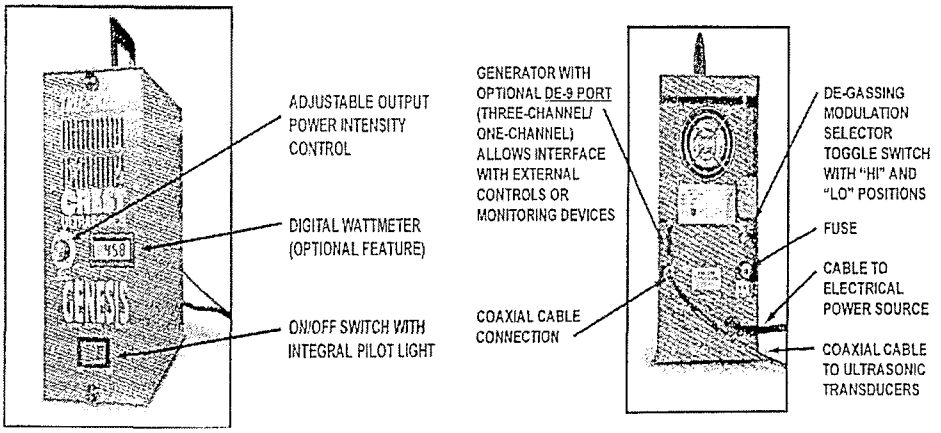

Figure 3.3: Genesis ultrasonic generator (front & rear) with adjustable output power intensity control (Crest Ultrasonic Product Bulletin, 2004).

3.2.1.2 Ultrasonic Transducer

Transducers were used in this experimental works for converting the signal received from ultrasonic generator into mechanical force. Three immiscible ultrasonic transducers '*piezoelectric type*' that supplied by Crest Ultrasonic (M) Sdn Bhd with frequencies 25 kHz used in this research. The transducers were selected because of its lower cost and easy availability (Crest Ultrasonic Product Bulletin, 2004).

Figure 3.4 shows the ultrasonic transducers used, where it utilized the piezoelectric effect because magnetostrictive transducers are generally less efficient than piezoelectric. Magnetostrictive transducer requires a dual energy conversion from electrical to magnetic and then from magnetic to mechanical. Some efficiency is lost in each conversion. Piezoelectric transducers convert alternating electrical energy directly to mechanical energy using piezoelectric effect. Table 3.1 summarizes the type and specification of the ultrasonic transducers used.

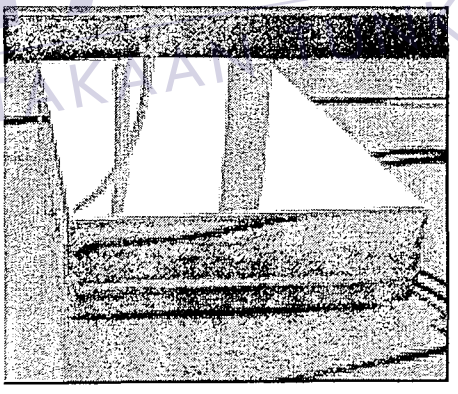

Figure 3.4: Immiscible ultrasonic transducer

Table 3.1: Description of ultrasonic transducers

Transducer	Specifications
25 kHz	Material : 316 Stainless Steel Dimension: 8.5"(L)x13"(W)x4"(H) Tru Sweep:+- 1 kHz Mounting type: Flexi cable

3.3 Binary Mixture

The binary mixture studied in this study was methanol-methyl acetate. Table 3.2 listed the characteristics of the selected binary mixture.

Table 3.2: Characteristics of binary mixtures

Binary Mixture	Type of Mixture	Azeotropic Point
Methanol-Methyl Acetate	Non-polar	34.78 %mole Methanol

3.4 Material and Equipment

Table 3.3 shows the chemicals used in this experiment. No further purification needed before using these chemicals. Since batch distillation method is used to obtain the VLE data, the samples prepared must have different compositions varied from 0 to 100 % (*Modul Amali Makmal Unit Operasi 2*, 2004). Table 3.4 shows the variation of feed samples.

Table 3.3: List of chemicals

Chemicals	Purity, %	Formula	Supplier
Methanol	99.8	CH_3OH	R & M Chemicals, UK
Methyl Acetate	99.7	$C_3H_6O_2$	R & M Chemicals, UK

Table 3.4: List of feed samples

Samples	Methanol, mL	Methyl Acetate, mL	% mole methanol
S_1	231	19	95.98
S_2	162	88	78.33
S_3	102	148	57.51
S_4	53.4	196.6	34.78
S_5	25	225	17.91

3.5 Experimental Work

The experimental work was conducted after completing the equipment set up and preparing the binary mixtures. This laboratory work was about to obtain vapor-liquid equilibrium data for different concentration of potassium chloride in ultrasonic distillation system. The purpose of this experiment was to get the optimum concentration of the salt in order to separate the azeotropic mixture. The ultrasonic frequency used was 25 kHz while the intensity was 200 W/A.cm^2. All the experimental studies were unable to repeat because of time limitation. Only some of them were repeated and the average of the data was recorded.

The final part of this study was the data analysis. Analysis method was by using Perkin Elmer Gas Chromatography which used to measure liquid and vapor composition. VLE data obtained from the experimental studies was plotted as *xy* diagrams. The salt optimum concentration was determined by the VLE result obtained. Table 3.5 shows the summary of the parameters studied in this research.

Table 3.5: Summary of the parameters

Parameters	Conditions
1) Binary mixtures used	Methanol-Methyl Acetate
2) Frequency (kHz)	25
3) Intensity (W/A.cm^2)	200
3) Operating Pressure (atm)	Atmospheric pressure, 1

3.5.1 VLE Study of the salt concentration

This part discussed about the procedure to obtain vapor-liquid equilibrium for different composition of binary mixture with the presence of ultrasonic waves and salt under atmospheric condition. Firstly, a mixture containing 231 mL methanol and 19 mL methyl acetate (Sample S_1) was prepared and charged into the distillation flask. Then amount of potassium chloride that has been weighted was added in the mixture. Then, the sample was boiled until reached the equilibrium. Equilibrium was reached once the liquid and vapor temperature remained constant for about 10-20 minutes. When the equilibrium was reached, the samples from distillation flask and vapor-condensed phase were collected for further analysis. Syringe was used to suck out the liquid sample from the distillation flask. This experiment was repeated using samples listed in Table 3.4. Figure 3.5 shows the flow chart of the VLE experimental procedure.

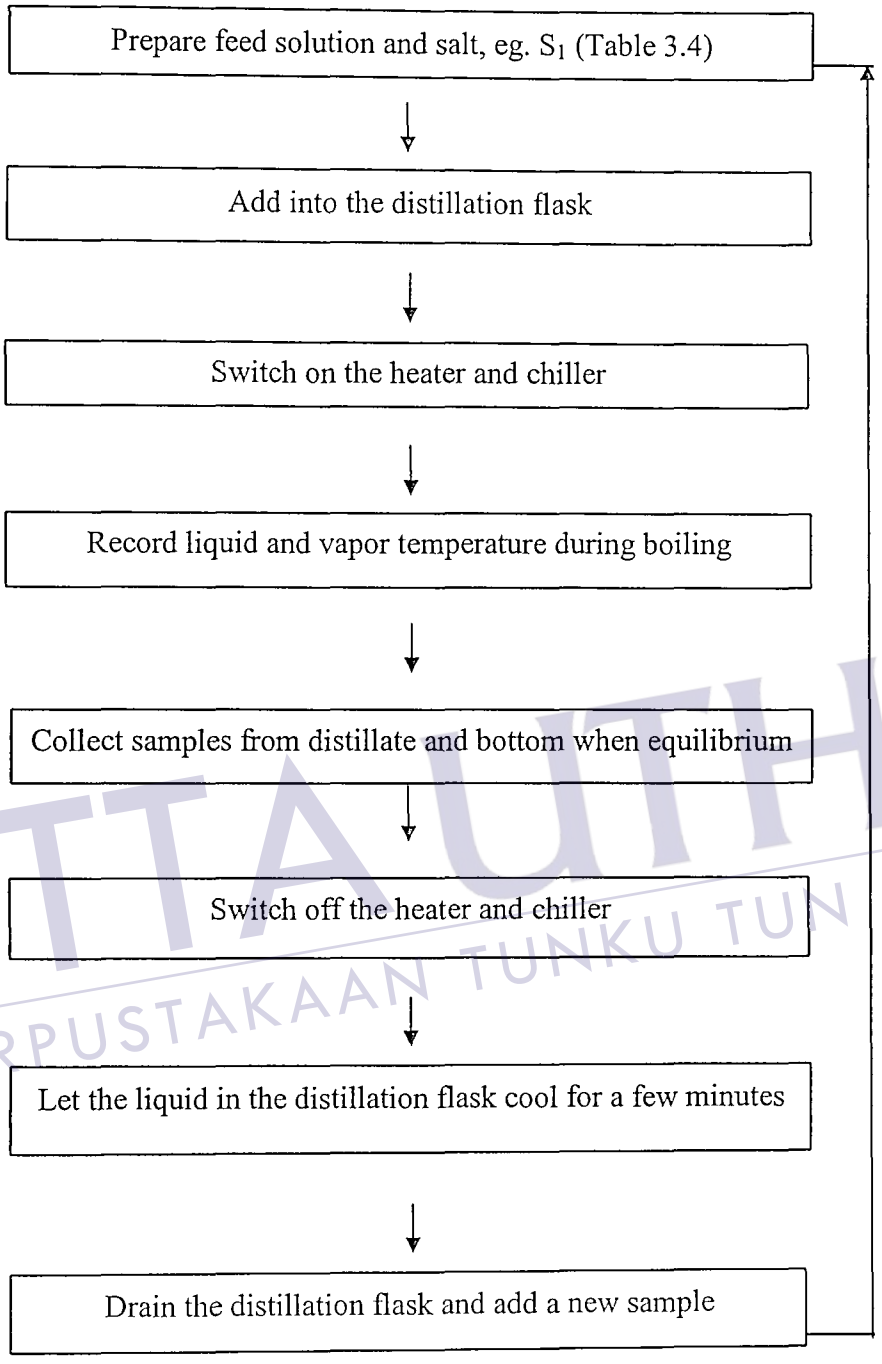

Figure 3.5: Flow chart of VLE experimental procedure

3.6 Analysis Method

The most important part in this research was the analysis of the VLE obtained from the experiment. The purpose of this part was for determining the composition of condensed vapor obtained which was methanol-methyl acetate. Gas chromatography (GC) is frequently used for the determination of the composition in liquid and condensed vapor phases of organic mixtures. The sample used in GC must either be in gas or capable of being converted to a gas at temperature of the column. Other methods for composition determination, although seldom applied in VLE measurements, include mass spectrometry, various spectroscopic methods, and measurements of density and refractive indices.

In this works, Perkin Elmer *Auto XL* Gas Chromatography System equipped with SPB-1 capillary column and FID detector as shows in Figure 3.5 was used. The analysis using gas chromatography was run under specific conditions as stated in Table 3.6.

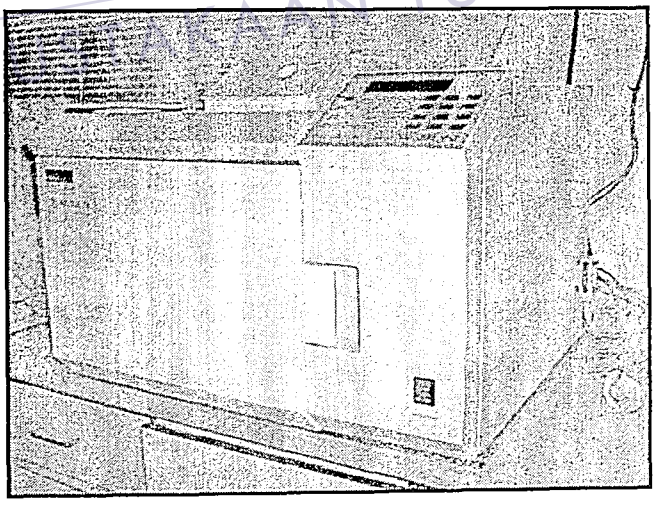

Figure 3.6: Perkin Elmer *Auto XL* Gas Chromatography System

Table 3.6: Analysis specification using gas chromatography

Samples Analyzed	GC Conditions
Methanol-Methyl Acetate	Column - SPB-1, 30m x 0.53 mm ID, 5.0 μm film Oven Temperature - 40 °C (5 min) to 100 °C at 5 °C/min Carrier Gas - Nitrogen 90 psi Detection - Flame Ionize Detection, 220 °C Injection Volume - 0.2 μL, split injection at 220 °C

Usually the quantitative measurement of a mixture performed using standard curve or working- curve method. The curve was obtained by plotting the peak area against concentration. The concentration is normally expressed as molar or weight percentage. In this research, the composition of the liquids and condensed vapors were determined by interpolating the plotted standard curve.

3.6.1 Method to Obtain Standard Curve

The standard curve can be obtained by preparing 10 mL mixtures of methanol and methyl acetate ranging from 0.0 to 1.0 mole fraction of methanol. Gas chromatography was used to measure the composition of each standard and also pure methanol. The area under the chromatographic peak was recorded and plotted against the compositions of solutions expressed in mole fraction of methanol. The composition of methanol in each sample collected from boiler and distillate was then measured using GC. The composition of the samples can be determined by interpolating the plotted graph. Table 3.7 depicts the standard mixtures of methanol-methyl acetate.

Table 3.7: Standard mixtures of Methanol-Methyl Acetate

Methanol Volume, mL	0.0	1.0	3.0	5.0	7.0	9.0	10.0
Methyl Acetate Volume, mL	10.0	9.0	7.0	5.0	3.0	1.0	0.0
Mole fraction Methanol, %	0	17.9	45.68	54.09	82.1	94.64	100

3.6.2 Standard Curves of Binary Mixture

Figure 3.6 shows the standard curves for binary mixtures of methanol-methyl acetate. This curve used for determining the composition of the collected samples.

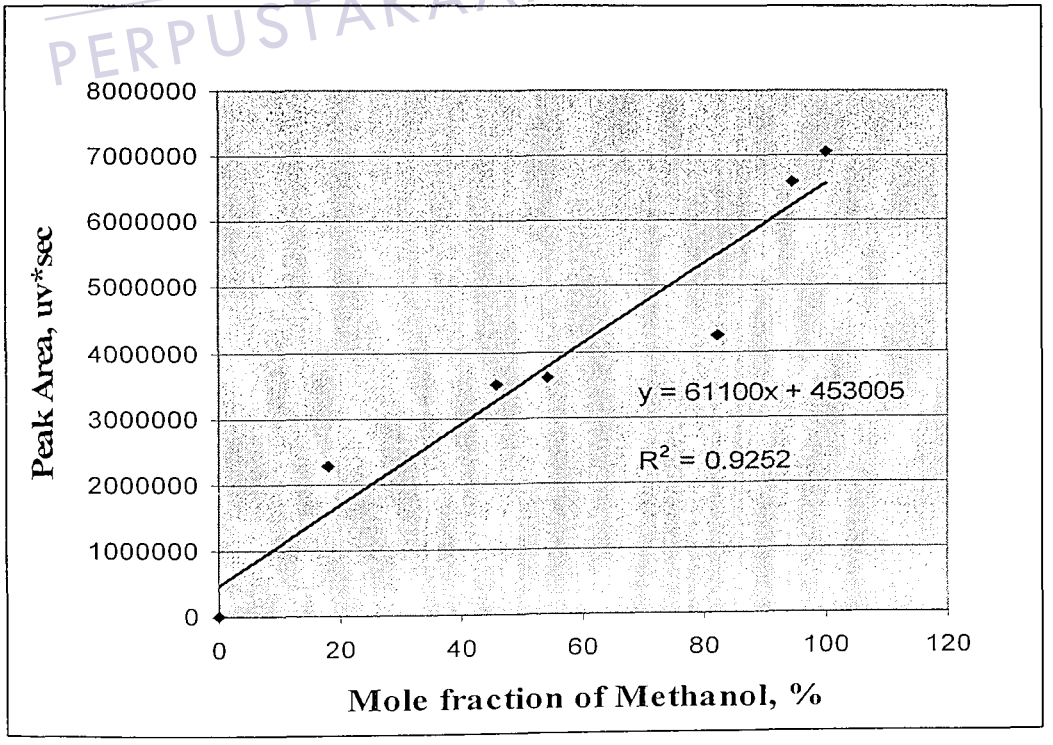

Figure 3.7: Peak area against mole fraction of Methanol

CHAPTER IV

RESULTS AND DISCUSSION

4.1 Introduction

This chapter discussed the results obtained from the experimental work. Only one type of mixture was studied as a sample. The mixture was methanol-methyl acetate where it forms an azeotropic mixture that is impossible to separate by using conventional distillation column. Thus, this study was conducted by using ultrasonic distillation column with salt as the separating agent in order to see the salt's effect on the separation process. The indicator parameters used to measure the salt effect to ultrasonic distillation system were equilibrium curve, azeotropic point and relative volatility. These parameters will be discussed in details. The VLE results obtained from VLE experimental studies with salt were also compared with VLE data obtained from experimental studies without salt.

4.2 Vapor-Liquid Equilibrium Graphs

The experiment was started by run the distillation without the presence of potassium chloride. This was done in order to investigate whether the use of salt give any differences in the separation process. In this condition, the separation only influenced by the ultrasonic wave that supplied with the frequency of 25 kHz and intensity of 200 W/A.cm^2. Figure 4.1 depicts the VLE without any presence of salt, Co (0 wt%). Then, the experiments were proceeding by adding potassium chloride in the mixture with different concentrations. Figure 4.2, 4.3 and 4.4 show the VLE graph with the presence of salt but with different concentrations, C1 (5 wt%), C2 (10 wt%) and C3 (15 wt%).

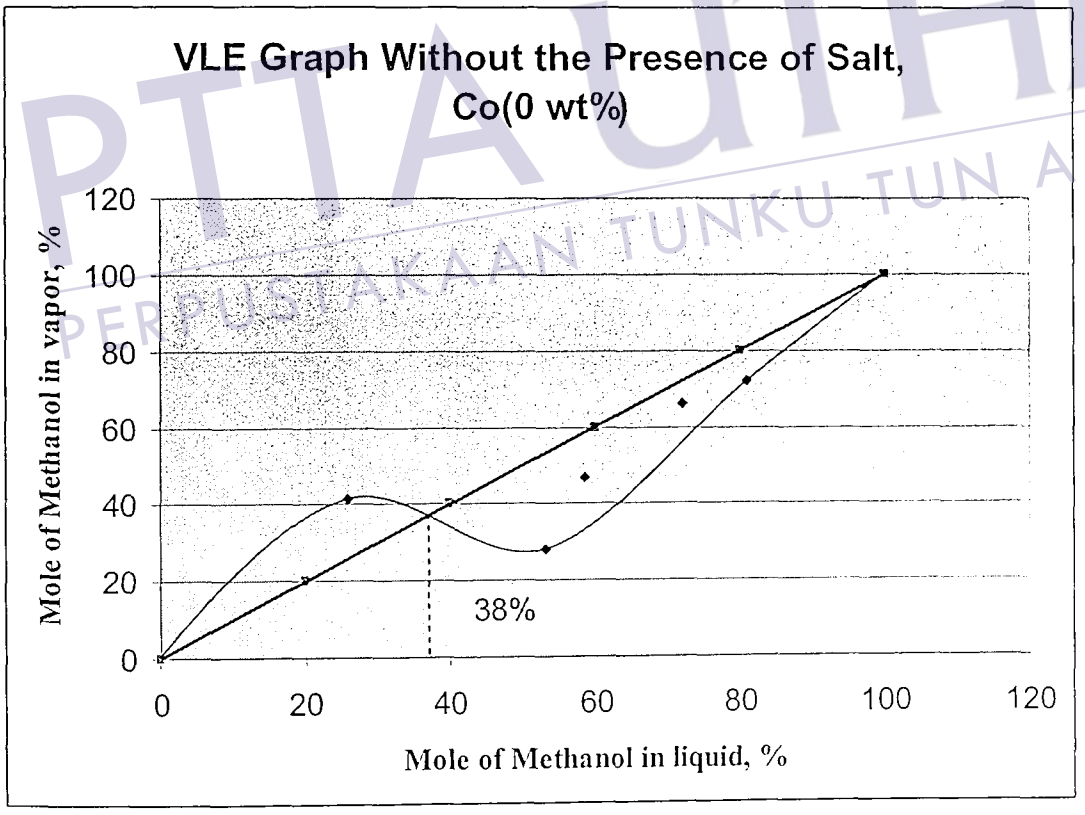

Figure 4.1: *xy*-diagram of methanol-methyl acetate system without the presence of salt, Co(0 wt%)

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